**Rules for oxidation numbers**

1) For any element in the ground state (the state that it is naturally in) the oxidation number = zero.

2) The oxidation numbers are assigned in this order F = -1, O = -2, H = +1.

3) For monoatomic ions (ions that have only one element in them) the oxidation number = charge.

4) For polyatomic ions (ions that have more than one element in them) the SUM of the oxidation numbers has to equal the charge.

5) For compounds, the SUM of the oxidation numbers has to equal zero.

Find the oxidation numbers for each element in:

Cl1- MnO41- Zn2+

C2O42- Fe3+ ClO41-

 NO31- SO3 SO42-